Chemistry 30: Hess’s Law pre- Lab

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_ Score \_\_\_\_\_\_ / 5

1. NaOH(s) is deliquescent. What does this mean?
2. Name the two processes are you putting the NaOH(s) through in this lab. \_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_
3. This lab has no instructions about making the initial temperature of the water used come to some predefined temperature. Why do you think it is all right just to take the water out of the tap and go with it?
4. This reaction is a classic acid base neutralization. For any acid or base, fill in the generic answers for products

Acid + Base 🡪 \_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. If the products have less potential energy than the reactants, then the reaction will \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (gain or release) thermal energy.