Chemistry 20: Five Day Copper Recovery Lab

Teacher NOTES

A prelab quiz each day keeps the students up to date. Although this is a 5 day lab, some days will only take 15 minutes or so. It is also possible to more than one step per day if time is an issue

The number of significant digits that a student can measure will determine what the rounding should be for all calculations.

I have assumed students have access to scales measuring to the nearest hundredth 🡪 so three significant digits

Objective: Recover 1.50 g of Cu(s) after making 5 chemical conversions

Materials: Copper wire

 Assorted beakers

 Dilute nitric acid

 Watch glass

 pH paper

 NaOH(aq)

 Ice/snow bath

 Distilled Water

 Dilute sulfuric acid

 Zinc metal

***Conversion 1*  Change Cu(s) to Cu(NO3)2(aq)**

* Take approximately 1.50  g of copper wire. Clean off any corrosion using steel wool
* You may find it useful to cut the copper into small pieces as this will facilitate a quicker reaction.
* Find the exact mass and record this value \_\_\_\_\_
* Place the copper into a clean 400 mL beaker. Carefully add 20 mL of dilute nitric acid
* Place a watch glass on top and observe the reaction that is taking place. Record your observations.
* Be patient as this is a SLOW reaction. Do not expect immediate changes.
* Then place the beaker in the fume hood until the copper has completely dissolved
	+ The NO2(g) is NOT healthy … so use some caution here. Overnight the fumes will dissipate … but don’t let the students breathe in much of this stuff!
* The brownish-orange gas produced by the reaction is NO2(g). The blue colour of the solution is characteristic of many copper compounds dissolved in water.
* Write the non, total and net ionic equations for this conversion.
	+ - Copper + nitric acid 🡪 copper (II) nitrate + nitrogen dioxide + water

Non: Cu(s) + 4HNO3(aq) 🡪 Cu(NO3)2(aq) + 2NO2(g) + 2 H2O(l)

Total: Cu(s) + 4H+(aq) + 4NO3-(aq) 🡪 Cu2+(aq) + 2NO3-(aq) + 2NO2(g) + 2H2O(l)

Net Cu(s) + 4H+(aq) + 2NO3-(aq) 🡪 Cu2+(aq) + 2NO2(g) + 2H2O(l)

**Calculate the volume of NO2(g) produced at SATP**

**Moles of Cu x x 22.4  = 1.06 L**

**Calculate the mass of Cu(NO3)2(aq) that will be dissolved in the solution.**

**Moles of Cu x  x 187.57 = 4.43 g**

**What has caused the solution to be blue in color?**

**Cu2+(aq) ions are blue in colour.**

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***Conversion 2:***  **Change Cu(NO3)2(aq) to Cu(OH)2(s)**

* Test the solution from conversion 1 with pH paper. Record the pH value
	+ Should be low as acid was in excess in conversion 1
* Now obtain 20 mL of NaOH (aq). Test this solution with pH paper. Record the pH value
	+ Should test high as this is a concentrated base
* Fill a large beaker one third full of ice/snow and water. Carefully place the beaker containing conversion 1 inside. (like a double boiler)
* Cautiously and carefully add the 20 mL of NaOH(aq) to conversion 1.
* Mix the solutions with a gentle swirling motion. The NaOH(aq) neutralizes the excess acid from conversion 1.
	+ This is an exothermic reaction and the beakers may become hot. BE CAREFUL.
* Test the pH of the mixture again and record \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
* If the pH is not as high as the pH of the NaOH(aq), keep adding more NaOH(aq) until it does match.
	+ You want the NaOH(aq) to become the EXCESS reagent
* A pale blue solid precipitate should have formed.
* Write the non, total and net ionic equations for this conversion.
* Copper(II) nitrate + sodium hydroxide 🡪 copper (II) hydroxide + sodium nitrate
* Cu(NO3)2(aq) +2 NaOH(aq) 🡪 Cu(OH)2(s) + 2NaNO3(aq)
* Cu2+(aq) + 2NO3-(aq) + 2Na+(aq) + 2OH-(aq) 🡪 Cu(OH)2(s) + 2Na+(aq) + 2NO3-(aq)
* Cu2+(aq) + 2OH-(aq) 🡪 Cu(OH)2(s)
* Calculate the mass of Cu(OH)2(s) that is produced in this conversion

x  x 97.57 ** = 2.32 g**

* What is a precipitate?

A precipitate is an ionic compound that is low in solubility. It will remain as an undissolved solid in the solution. Remember that dissolving is temperature dependent.

* Why did you continue to add the NaOH(aq) until the pH became high?

To ensure that all the copper (II) nitrate is actually converted to copper (II) hydroxide, the excess acid must be completely used up and the sodium hydroxide must be in excess.

***Conversion 3:***  **Change Cu(OH)2(s) to CuO(s)**

* Add 100 mL of distilled water to the beaker containing conversion 2
* Using a hotplate and eye protection, carefully heat the beaker until the solution begins to boil. STIR!
	+ This is the step that students tend to have trouble with. If they get the heat too high, the solution will “spit” chunks out and they will lose some of their copper.
	it goes without saying that EYE protection is a MUST until all the students are done heating glassware!
* A brown-black precipitate will form
* Be careful that the solution does not spit out of the beaker or you will lose part of your copper for recovery!
* Remove from the heat and let the solution cool for at least 5 minutes
* Use a wash bottle to rinse the stir rod back into the beaker
* Decant the extra liquid. Ensure that no solid is lost.
* Wash the precipitate by adding an additional 100 mL of distilled water. STIR
* Let the precipitate settle again and decant off the extra liquid.
* Write the non, total and net ionic reactions for this conversion.
* Copper(II) hydroxide 🡪 copper (II) oxide + water
* Cu(OH)(s) 🡪 CuO(s) + H2O(l)
* Cu(OH)(s) 🡪 CuO(s) + H2O(l)
* Cu(OH)(s) 🡪 CuO(s) + H2O(l)

Name the black solid produced in this step copper (II) oxide

Calculate the mass of copper (II) oxide produced in this conversion

x  x 79.55 ** = 1.90 g**

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***Conversion 4:***  **Change CuO(s) to CuSO4(aq)**

* Add 50 mL of dilute sulfuric acid to the black copper (II) oxide.
* Stir gently. The copper (II) oxide will dissolve and the solution will turn blue again.
* Write the non, total and net ionic equations for this conversion.
* Copper(II) oxide + sulfuric acid 🡪 copper (II) sulfate + water
* CuO(s) + H2SO4(aq) 🡪 CuSO4(aq) + H2O(l)
* CuO(s) + 2H+(aq) + SO42-(aq) 🡪 Cu2+(aq) + SO42-(aq) + H2O(l)
* CuO(s) + 2H+(aq) 🡪 Cu2+(aq) + H2O(l)

Determine the mass of copper (II) sulfate formed in this conversion

 x  x 159.62 ** = 3.79 g**

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***Conversion 5:***  **Change CuSO4(aq) to Cu(s)**

* Weigh about 7.0 g of Zn(s) and record \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	+ Zn(s) must be in excess. This will work better if the pieces of Zn(s) are not too big. It may give trouble later on if the pieces are really large.
* Add this Zn(s) to the copper (II) sulfate solution
* Immediately cover the beaker with a watch glass and allow it to stand until the blue colour disappears. You will have to swirl the beaker. (Be patient – this make take 10 – 15 minutes)
* If the reaction appears to slow or stop, add an additional 20 mL of acid.
* Let the beaker stand until the solid zinc metal disappears as well. Since the reaction involves the evolution of hydrogen gas, the **absence of bubbles** forming will be an indication that the zinc metal is all consumed
	+ Sometimes I have had the students let this sit overnight …. Depending on the size of the zinc chunks, this can be a SLOW process.
* You should now have the solid copper returning.!!
* There are actually two reactions that have happened!
* **The first explains the formation of the copper**
* Write a non, total and net ionic equation for this process.
* Copper(II) sulfate + Zinc 🡪 copper + Zinc sulfate
* CuSO4(aq) + Zn(s) 🡪 Cu(s) + ZnSO4(aq)
* Cu2+(aq) + SO42-(aq) + Zn(s) 🡪 Cu(s) + Zn2+(aq) + SO42-(aq)
* Cu2+(aq) + Zn(s) 🡪 Cu(s) + Zn2+(aq)
* **The second reaction explains how you get rid of the excess zinc**
* Write a non, total and net ionic equation for this process.
* Sulfuric Acid + Zinc 🡪 hydrogen gas + zinc sulfate
* H2SO4(aq) + Zn(s) 🡪 H2(g) + ZnSO4(aq)
* 2H+(aq) + SO42-(aq) + Zn(s) 🡪 H2(g) + Zn2+(aq) + SO42-(aq)
* 2H+(aq) + Zn(s) 🡪 H2(g) + Zn2+(aq)
* Now you need to collect and find the mass of the copper you have recovered.
	+ Allow the solid copper to settle to the bottom.
	+ Decant and discard the clear liquid
	+ Wash the solid copper a **minimum of three times**. Use 50 mL portions of hot (almost boiling) distilled water

This is important because if there are impurities left in the sample, it will darken and possible turn black again as it dries. A good sample should be the ‘coppery’ colour of pure copper.

* + Each time, stir, and then decant the liquid
	+ Accurately weigh an evaporating dish \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	+ Transfer the copper metal to the dish. Decant off as much excess water as possible.
		- Have them spread the copper sample out so it will dry faster. Remember that clothes in a ‘heap’ do not dry well. Neither will the sample!
	+ Allow the dish to sit and dry overnight (or several nights!).
	+ Find the mass of the dish + the copper \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	+ If the copper metal gets too hot, it will go back to copper (II) oxide!. This is a bad thing …… you have already done this step!

**Observations:**

|  |  |
| --- | --- |
| Conversion I |  |
| Conversion II |  |
| Conversion III |  |
| Conversion IV |  |
| Conversion V |  |

**Analysis of data**

Initial mass of copper \_\_\_\_\_\_\_\_\_\_\_

Final mass of copper and evaporating dish \_\_\_\_\_\_\_\_\_\_\_

Mass of evaporating dish \_\_\_\_\_\_\_\_\_\_\_

Mass of recovered copper \_\_\_\_\_\_\_\_\_\_\_

Use the mass of copper recovered to calculate the % recovery for this lab

Be sure to allow for sufficient time for the sample to dry or the students will all be OVER by a great amount.

Remember that there may be impurities trapped in their final sample

But in the same breath, the copper they started with may not have been pure either

Justify why you recover is

* Less than 100%
* More than 100%
* Exactly 100%

Choose only one of the above!!

This is a really cool lab …. Teaches lots of good techniques and reinforces that the limiting reagent is in charge of the reaction

It also gives them practice at ‘removing’ the excess reagents!